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Chemistry, 2012 EC, Grade 8 model exam

Grade 8 · Chemistry · 2012 E.C.

Time allowed: 120 minutes **Questions:** 58 **Total marks:** 58

Name _____ Grade _____ ID No. _____ Date _____

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1. The essence of chemistry concerned mainly with (1)

- **A.** Searching for solution to problem making
- **B.** Relying only an experimental proved facts
- **C.** Production of new materials of desirable properties
- **D.** All

2. In Ethiopia caustic soda is manufactured in (1)

- **A.** Addis Ababa
- **B.** Zeway
- **C.** Wenji
- **D.** Awash Melkasa

3. Which properties of substance is measurable? (1)

- **A.** color
- **B.** odor
- **C.** taste
- **D.** density

4. A chemical change is characterized by:- (1)

- **A.** energy change is involved
- **B.** no change in composition
- **C.** no change in the mass of the substance is involved
- **D.** the change is easy to reverse

5. What is the formula name of CU_2O? (1)

- **A.** Copper (II) oxide
- **B.** copper peroxide
- **C.** Copper (I) oxide
- **D.** copper (III) oxide

6. What does “X” represent in the following reaction? $Al_2O_3 + 2NaOH \rightarrow X + H_2O$ (1)

- **A.** $Al(OH)_3$
- **B.** AlH_3
- **C.** $Al_2O_3 \cdot H_2O$
- **D.** $NaAlO_2$

7. Substances that take part in a chemical reaction is called? (1)

- **A.** subscript
- **B.** products
- **C.** reactants
- **D.** None of the above

8. What is the chemical formula of Barium phosphate? (1)

- **A.** $Ba\,(PO_4)_2$
- **B.** $BaPO_4$
- **C.** $Ba_3\,(PO_4)_2$
- **D.** Ba_2PO_4

9. The liquid which passes through the filter paper is? (1)

A. Filter paper
B. funnel
C. residue
D. filtrate

10. Which one of the following technique is used to get a cup of clear coffee as it is poured from coffee Pot (jebena)? (1)

A. Filtration
B. decantation
C. distillation
D. evaporation

11. A block of ion metal occupies a volume of $20\ \mathrm{cm}^3$ and has a density of $7.86\ \mathrm{g/cm}^3$. What is the mass Of iron? (1)

A. 1572g
B. 157.2g
C. 15.72g
D. 1.752g

12. The coefficient "3" in the formula $3\mathrm{P}_4$ represents:- (1)

A. 3 atoms phosphorous
B. 4 molecules of phosphorous
C. 3 molecules of phosphorous
D. 12 molecules of phosphorous

13. Which of the following is the only liquid metallic element? (1)

A. Bromine
B. chlorine
C. mercury
D. hydrogen

14. Which one of the following is TRUE about properties of acid? (1)

A. acids have bitter taste
B. acids change the color of phenaophitaline into colorless
C. acids has a PH value greater than seven(7)
D. acids changes red litmus in to blue

15. Which one of the following is not a metalloid? (1)

A. silicon
B. arsenic
C. bromine
D. boron

16. Magnesium has a valence number of "2" and nitrogen has a valence number of "3" what is the Chemical formula for Magnesium nitride? (1)

A. $\mathrm{Mg_2N_3}$
B. $\mathrm{Mg_3N}$
C. $\mathrm{Mg_3N_2}$
D. MgN

17. Which one of the following is not homogeneous mixture? (1)

A. sea water
B. soft drink
C. brass
D. dusty air

18. Which pair of liquids are immiscible liquid ? (1)

A. vinegar and water
B. oil and water
C. water and alcohol
D. none of the above

19. A pure substance composed of two or more elements that are combined chemically in a definite Proportion by mass is called ..... (1)

A. element
B. mixture
C. compound
D. none

20. The temperature at which gas (vapor) directly change into solid is called? (1)

A. sublimation
B. evaporation
C. deposition
D. condensation

21. Which of the following apparatus is used for the separation of salt from salt solution? (1)

A. evaporating dish
B. filter paper
C. condenser
D. thermometer

22. What is the name of an element having the chemical symbol “Sn”? (1)

A. tin
B. sulphur
C. sodium
D. silver

23. The combining power of an element is called? (1)

A. Valence electron
B. valence number
C. valence shell
D. coefficient

24. Which of the following is not poly atomic molecule? (1)

A. chlorine
B. ozone
C. phosphorous
D. sulphur

25. In balancing a chemical equation using the least common multiple (LCM) method the following steps Are followed. I. Find the LCM of the total valence number and write it on the arrow. II. Write the word equation. III. Write the total valence number of each of the elements above their symbol. Iv. Change the word equation to formula equation. v. divide the LCM by each total valence number and use dividend as the coefficient for each species in the equation. Which of the following is the correct step for balancing a chemical equation using the LCM methods? (1)

A. I, III, II, IV, V
B. II,I,III,IV,V
C. II,IV,III,I,V
D. III,II,IV,I,V

26. The atomic masses of the two stable isotopes ofboron, $^{10}\mathrm{B}$(20%) and $^{11}\mathrm{B}$(80%) are 10 and 11 Respectively. What is the average atomic mass of boron? (1)

A. 108 amu
B. 1.08amu
C. 10.8amu
D. 0.108amu

27. The correct balanced chemical equation for the reaction between lead and oxygen to form lead(II)oxide is (1)

A. $\mathrm{Pb} + \mathrm{O_2} \rightarrow \mathrm{PbO}$
B. $2\mathrm{Pb} + \mathrm{O_2} \rightarrow 2\mathrm{PbO}$
C. $4\mathrm{Pb} + \mathrm{O_2} \rightarrow 2\mathrm{Pb_2O}$
D. $\mathrm{Pb} + \mathrm{O_2} \rightarrow \mathrm{PbO_2}$

28. Which of the following atomic energy level is nearest to the nucleus of an atom? (1)

A. L-shell
B. K-Shell
C. M-shell
D. N-shell

29. Which of the following is not molecules of compound? (1)

A. $\mathrm{O_3}$
B. $\mathrm{H_2O}$
C. $\mathrm{P_4}$
D. except “B”

30. Which of the following statements describes the shortcoming of Dalton atomic theories? (1)

A. Atoms combine in a fixed small whole number ratio to form compounds.

B. Atoms of different elements are different

C. Atoms of the elements have the mass

D. Matter consists of very small particles called atoms.

31. The horizontal rows of elements in the periodic table are called:- (1)

A. Group

B. Coefficient

C. period

D. Subscript

32. Which one of the following subatomic particles has the smallest mass? (1)

A. Proton

B. Neutron

C. Electron

D. Nucleus

33. How many electrons does O^{-2} have? (1)

A. 10

B. 8

C. 6

D. 2

34. What are the number of proton, electron and neutron in $^{24}_{12}Mg^{+2}$ respectively? (1)

A. 12.10,12

B. 12, 14, 12

C. 12,12,12

D. 14,12,12

35. An element with atomic number 23 is found in ..... (1)

A. group IIIA and period 4

B. group VA and period 4

C. group IVA and period 4.

D. group 4 and period V

36. The atomic number of chromium is 24.How many electrons are contained in its third energy level? (1)

A. 18

B. 8

C. 32

D. 2

37. The sum of protons and neutrons in an atom is called; (1)

A. Atomic number

B. Mass number

C. Atomic mass

D. None of the above

38. The common name of group VIIIA element is..... (1)

A. Noble gas

B. Chalcogen

C. Halogens

D. Alkaline earth metals

39. Across a period atomic size of the elements..... (1)

A. Increase

B. remains the same

C. Decrease

D. Increase and decrease

40. Which one of the following is inorganic compound? (1)

A. animal fat

B. sugar

C. vinegar

D. carbon dioxide

41. Which of the following hydrocarbon used to produce oxyacetylene flam? (1)

A. Octane
B. Decane
C. Ethene
D. Ethyne

42. The mixture of carbon monoxide and hydrogen gas is produce ..... (1)

A. water gas
B. producer gas
C. biogas
D. methane

43. Which one is the correct general formula for Alkyne? (1)

A. C_nH_{2n}
B. C_nH_{2n-2}
C. C_nH_{2n+2}
D. C_nH_{2n-1}

44. Which of the following oxide is binary compound? (1)

A. CO_2
B. MgO
C. CrO_3
D. All of them

45. The characteristics property of base is due to the presence of :- (1)

A. Hydrogen ion
B. Hydroxide ion
C. Oxide ion
D. Hydride ion

46. Which of the following salts is used as a baking powder? (1)

A. $CaCO_3$
B. $NaHCO_3$
C. $(NH_4)_2HPO_4$
D. KNO_3

47. Carbonates are salts of:- (1)

A. Carbonic acid
B. Sulphuric acid
C. Phosphoric acid
D. Hydrochloric acid

48. Which of the following pairs of compounds produces sodium chloride? (1)

A. Sodium Oxide and water
B. Sodium oxide and sulphuric acid
C. Sodium Hydroxide and hydrochloric acid
D. Sodium Hydroxide and Nitric acid

49. What measure we talk, if you drink corrosive acid? You are advisable to drink:- (1)

A. KOH
B. $NaOH$
C. $Mg(OH)_2$
D. $Ca(OH)_2$

50. Which gas is given offby the reaction between acid and carbonates? (1)

A. CO_2
B. O_2
C. H_2
D. SO_2

51. Which one of the following is not the properties of metal? (1)

A. Luster
B. Malleable
C. sonorous
D. Metals have slightly less melting and boiling point

52. Which one is the seventh most abundant element and the fifth most abundant metal in the EarthCrust? (1)

A. Potassium
B. Sodium
C. Magnesium
D. Aluminum

53. Which of the following are does not contain copper? (1)

A. Malachite
B. chalcocite's
C. Magnetite
D. Coveline

54. Which of the following statements does not describe gold? (1)

A. it is an excellent conductor
B. it does not exists as free element in nature
C. It does not react with oxygen in the presence of moist air
D. it has a yellow color

55. Which of the following alloys does not contain copper? (1)

A. brass
B. bronze
C. Cupronickel
D. steenless steel

56. Which of the following elements does not have an allotropic forms? (1)

A. sulphur
B. phosphorus
C. carbon
D. Nitrogen

57. Which one is not the important ore of Aluminum? (1)

A. corundum
B. Bauxite
C. Sylvite
D. Cryoline

58. Which of the following Nonmetallic element are solid at room temperature? (1)

A. Fluorine
B. Bromine
C. Iodine
D. Chlorine

END OF THE EXAMINATION

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